Test 5

Dissolved Oxygen

INTRODUCTION

Oxygen gas dissolved in water is vital to the existence of most aquatic organisms. Oxygen is a key component in cellular respiration for both aquatic and terrestrial life. The concentration of dissolved oxygen, *DO*, in an aquatic environment is an important indicator of the environment's water quality.

Table 1: Minimum DO Requirements	
Organism	Minimum dissolved oxygen (mg/L)
Trout	6.5
Smallmouth bass	6.5
Caddisfly larvae	4.0
Mayfly larvae	4.0
Catfish	2.5
Carp	2.0
Mosquito larvae	1.0

Some organisms, such as salmon, mayflies, and trout, require high concentrations of dissolved oxygen. Other organisms, such as catfish, mosquito larvae, and carp, can survive in environments with lower concentrations of dissolved oxygen. The diversity of organisms is greatest at higher DO concentrations. Table 1 lists the minimum dissolved oxygen concentrations necessary to sustain selected animals.

Oxygen gas is dissolved in water by a variety of processes—diffusion between the atmosphere and water at its surface, aeration as water flows over rocks and other debris, churning of water by waves and wind, and photosynthesis of aquatic plants. There are many factors that affect the concentration of dissolved oxygen in an aquatic environment. These factors include: temperature, stream flow, air pressure, aquatic plants, decaying organic matter, and human activities.

Sources of DO

- Diffusion from atmosphere
- Aeration as water moves over rocks and debris
- Aeration from wind and waves
- Photosynthesis of aquatic plants

As a result of plant activity, DO levels may fluctuate during the day, rising throughout the morning and reaching a peak in the afternoon. At night photosynthesis ceases, but plants and animals continue to respire, causing a decrease in DO levels. Because large daily fluctuations are possible, DO tests should be performed at the same time each day. Large fluctuations in dissolved oxygen levels over a short period of time may be the result of an algal bloom. While the algae population is growing at a fast rate, dissolved oxygen levels increase. Soon the algae begin to die and are decomposed by aerobic bacteria, which use up the oxygen. As a greater number of algae die, the oxygen requirement of the aerobic decomposers increases, resulting in a sharp drop in dissolved oxygen levels. Following an algal bloom, oxygen levels can be so low that fish and other aquatic organisms suffocate and die.

Factors that affect DO levels

- Temperature
- Aquatic plant populations
- Decaying organic material in water
- Stream flow
- Altitude/atmospheric pressure
- Human activities

Temperature is important to the ability of oxygen to dissolve, because oxygen, like all gases, has different solubilities at different temperatures. Cooler waters have a greater capacity for dissolved oxygen than warmer waters. Human activities, such as the removal of foliage along a stream or the release of warm water used in industrial processes, can cause an increase in water temperature along a given stretch of the stream. This results in a lower dissolved oxygen capacity for the stream.

Expected Levels

The unit mg/L[2] is the quantity of oxygen gas dissolved in one liter of water. When relating DO measurements to minimum levels required by aquatic organisms, mg/L is used. The procedure described in this chapter covers the use of a Dissolved Oxygen Probe to measure the concentration of DO in mg/L. Dissolved oxygen concentrations can range from 0 to 15 mg/L. Cold mountain streams will likely have DO readings from 7 to 15 mg/L, depending on the water temperature and air pressure. In their lower reaches, rivers and streams can have DO readings between 2 and 11 mg/L.

Table 2	
DO Level	Percent Saturation of DO
Supersaturation[1]	≥ 101%
Excellent	90 – 100%
Adequate	80 – 89%
Acceptable	60 – 79%
Poor	< 60%

When discussing water quality of a stream or river, it can be helpful to use a different unit than mg/L. The term percent saturation is often used for water quality comparisons. Percent saturation is the dissolved oxygen reading in mg/L divided by the 100% dissolved oxygen value for water (at the same temperature and air pressure). The manner in which percent saturation relates to water quality is displayed in Table 2. In some cases, water can exceed 100% saturation and become supersaturated for short periods of time.

Summary of Methods

Dissolved oxygen can be measured directly at the site or from water samples transported from the site. Measurements can be made at the site by either placing the Dissolved Oxygen Probe directly into the stream away from the shore or by collecting a water sample with a container or cup and then taking measurements with the Dissolved Oxygen Probe back on the shore. Water samples collected from the site in capped bottles and transported back to the lab must be stored in an ice chest or refrigerator until measurements are to be made. Transporting samples is not recommended, because it reduces the accuracy of test results.

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1 Supersaturation can be harmful to aquatic organisms. It can result in a disease known as Gas Bubble Disease.

2 The unit of mg/L is numerically equal to parts per million, or ppm.

DISSOLVED OXYGEN

Materials Checklist

___ laptop computer (Power Mac or Windows)
___ Vernier computer interface, battery-powered
___ Logger *Pro*
___ Vernier Dissolved Oxygen Probe
___ DO Electrode Filling Solution
___ Sodium Sulfite Calibration Solution
___ 250-mL beaker
___ 100% calibration bottle
___ wash bottle with distilled water
___ tissues or paper towels
___ pipet
___ small plastic or paper cup (optional)

Collection and Storage of Samples

1. Before you begin sampling, fill out the site information on the Data & Calculations sheet. Space for observations regarding the site is provided at the bottom of the Data & Calculations sheet. Special things to note about the site are the weather, descriptions of the stream reach (flow, depth, shape), and a description of the riparian zone (density of foliage and width of riparian zone).

2. It is important to sample as far away from the shore as is safe and under the surface of the water. Samplers consisting of a rod and container can be constructed for collection of samples from areas of the stream otherwise unreachable. Refer to page Intro-4 of the Introduction of this book for more details. In slow-moving water, it is necessary to take samples below the water's surface at various depths.

3. When collecting a sample with a cup or container, prevent mixing of the water sample and air by collecting your sample from below the water surface.

4. If you are going to take readings after returning to the laboratory, make sure that there are no air bubbles in the water-sample container and that the container is tightly stoppered. The sample should be stored in an ice chest or refrigerator until measurements are to be made. Storing water samples for later testing decreases sample accuracy and is only recommended in cases where measuring at the site is not possible.

5. When taking readings in cold (0–10°C) or warm (25–35°C) water, allow more time for the dissolved oxygen readings to stabilize. Automatic temperature compensation in the Dissolved Oxygen Probe is not instantaneous and readings may take up to 2 minutes to stabilize depending on the temperature.

Testing Procedure

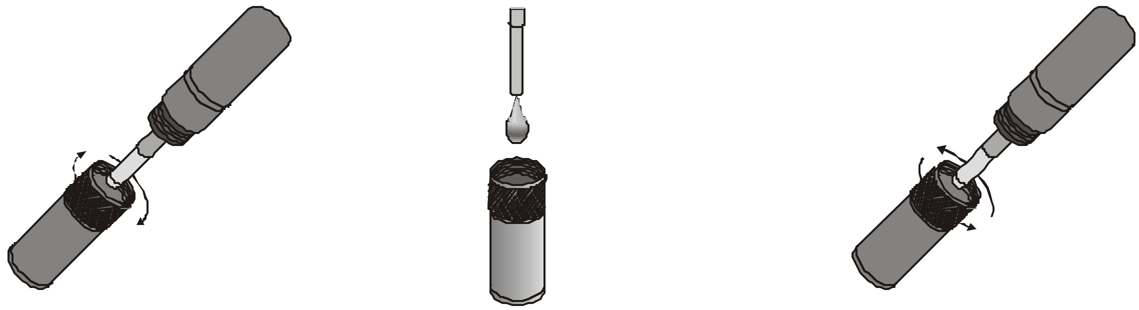
1. Position the computer safely away from the water. Keep water away from the computer at all times.

2. Prepare the computer for data collection by opening "Test 05 Dissolved Oxygen" from the *Water Quality with Computers* experiment files of Logger*Pro*. On the Graph window, the vertical axis has dissolved oxygen concentration scaled from 0 to 14 mg/L. The horizontal axis has time scaled from 0 to 10 seconds. There is also a Meter window which displays live dissolved oxygen concentration readings.

3. Plug the Dissolved Oxygen Probe into Port 1 or Channel 1 of the Vernier computer interface.

4. Prepare the Dissolved Oxygen Probe for use.

   a. Unscrew the membrane cap from the tip of the probe.
   b. Using a pipet, fill the membrane cap with 1 mL of DO Electrode Filling Solution.
   c. Carefully thread the membrane cap back onto the electrode.
   d. Place the probe into a 250-mL beaker containing distilled water.



Remove membrane cap — Add electrode filling solution — Replace membrane cap

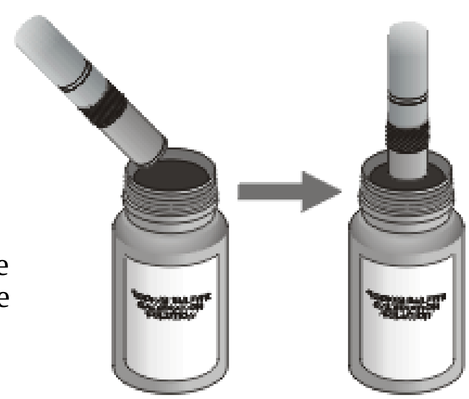
5. It is necessary to warm up the Dissolved Oxygen Probe for 30 minutes before taking readings. This warm-up period is known as *polarizing*. To polarize the probe, leave the probe connected to the interface, with Logger *Pro* running, for 30 minutes. The probe must stay connected at all times to keep it warmed up. If disconnected for a period longer than 10 minutes, it will be necessary to polarize the probe again.

6. You are now ready to calibrate the Dissolved Oxygen Probe.

   - If your instructor directs you to use the calibration stored in the experiment file, then proceed to Step 7.
   - If your instructor directs you to perform a new calibration for the Dissolved Oxygen Probe, follow this procedure.

   Zero-Oxygen Calibration Point

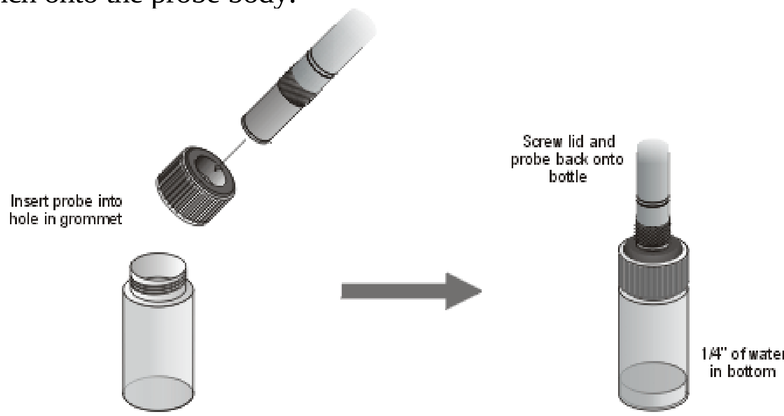
   a. Choose Calibrate from the Experiment menu and then click [Perform Now].
   b. Remove the probe from the water bath and place the tip of the probe into the Sodium Sulfite Calibration Solution. **Important:** No air bubbles can be trapped below the tip of the probe or the probe will sense an inaccurate dissolved oxygen level. If the voltage does not rapidly decrease, tap the side of the bottle with the probe to dislodge any bubbles. The readings should be in the 0.2- to 0.5-V range.
   c. Type “0” (the known value in mg/L) in the edit box.
   d. When the displayed voltage reading for Input 1 stabilizes, click [Keep].



Insert probe at an angle — Submerge probe tip 1-2 cm

Saturated DO Calibration Point

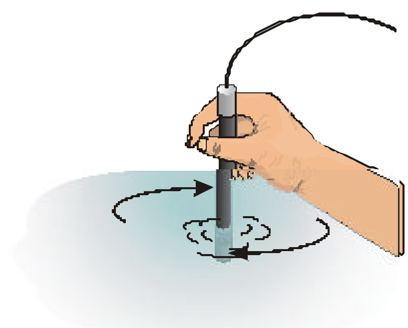
e. Rinse the probe with distilled water.
f. Unscrew the lid of the calibration bottle provided with the probe. Slide the lid and the grommet about 1/2 inch onto the probe body.



g. Add water to the bottle to a depth of about 1/4 inch and screw the bottle into the cap, as shown. **Important:** Do not touch the membrane or get it wet during this step. Keep the probe in this position for about a minute.
h. Type the correct saturated dissolved-oxygen value (in mg/L) from Table 3 (for example, "8.66") using the current barometric pressure and air temperature values. If you do not have the current air pressure, use Table 4 to estimate the air pressure at your altitude.
i. When the displayed voltage reading for Input 1 stabilizes (readings should be in the 2.5- to 3.0-V range), click Keep and then click OK.

7. You are now ready to collect dissolved oxygen concentration data.

   a. Rinse the tip of the probe with sample water.
   b. Place the tip of the probe into the stream at Site 1, or into a cup with sample water from the stream. Submerge the probe tip to a depth of 4-6 cm.
   c. Gently stir the probe in the water sample. Monitor the dissolved oxygen concentration value in the Meter window. **Note:** It is important to keep stirring until you have finished collecting DO data.
   d. If the DO value appears stable, simply record it on the Data & Calculations sheet and proceed to Step 9.



8. If the DO value displayed in the Meter window is fluctuating, determine the *mean* (or average) dissolved oxygen concentration. To do this:

   a. Click Collect to begin a 10-second sampling run. **Important:** Leave the probe tip submerged and continue to stir for the 10 seconds that data is being collected.
   b. Click on the Statistics button, STAT, to display the statistics box on the graph.
   c. Record the mean dissolved oxygen concentration value on the Data & Calculations sheet.

9. Return to Step 7 to obtain a second reading. When both readings have been taken, rinse the tip of the probe and secure it in the calibration bottle filled with water.

Table 3: 100% Dissolved Oxygen Capacity (mg/L)												
	770 mm	**760 mm**	**750 mm**	**740 mm**	**730 mm**	**720 mm**	**710 mm**	**700 mm**	**690 mm**	**680 mm**	**670 mm**	**660 mm**
0°C	14.76	14.57	14.38	14.19	13.99	13.80	13.61	13.42	13.23	13.04	12.84	12.65
1°C	14.38	14.19	14.00	13.82	13.63	13.44	13.26	13.07	12.88	12.70	12.51	12.32
2°C	14.01	13.82	13.64	13.46	13.28	13.10	12.92	12.73	12.55	12.37	12.19	12.01
3°C	13.65	13.47	13.29	13.12	12.94	12.76	12.59	12.41	12.23	12.05	11.88	11.70
4°C	13.31	13.13	12.96	12.79	12.61	12.44	12.27	12.10	11.92	11.75	11.58	11.40
5°C	12.97	12.81	12.64	12.47	12.30	12.13	11.96	11.80	11.63	11.46	11.29	11.12
6°C	12.66	12.49	12.33	12.16	12.00	11.83	11.67	11.51	11.34	11.18	11.01	10.85
7°C	12.35	12.19	12.03	11.87	11.71	11.55	11.39	11.23	11.07	10.91	10.75	10.59
8°C	12.05	11.90	11.74	11.58	11.43	11.27	11.11	10.96	10.80	10.65	10.49	10.33
9°C	11.77	11.62	11.46	11.31	11.16	11.01	10.85	10.70	10.55	10.39	10.24	10.09
10°C	11.50	11.35	11.20	11.05	10.90	10.75	10.60	10.45	10.30	10.15	10.00	9.86
11°C	11.24	11.09	10.94	10.80	10.65	10.51	10.36	10.21	10.07	9.92	9.78	9.63
12°C	10.98	10.84	10.70	10.56	10.41	10.27	10.13	9.99	9.84	9.70	9.56	9.41
13°C	10.74	10.60	10.46	10.32	10.18	10.04	9.90	9.77	9.63	9.49	9.35	9.21
14°C	10.51	10.37	10.24	10.10	9.96	9.83	9.69	9.55	9.42	9.28	9.14	9.01
15°C	10.29	10.15	10.02	9.88	9.75	9.62	9.48	9.35	9.22	9.08	8.95	8.82
16°C	10.07	9.94	9.81	9.68	9.55	9.42	9.29	9.15	9.02	8.89	8.76	8.63
17°C	9.86	9.74	9.61	9.48	9.35	9.22	9.10	8.97	8.84	8.71	8.58	8.45
18°C	9.67	9.54	9.41	9.29	9.16	9.04	8.91	8.79	8.66	8.54	8.41	8.28
19°C	9.47	9.35	9.23	9.11	8.98	8.86	8.74	8.61	8.49	8.37	8.24	8.12
20°C	9.29	9.17	9.05	8.93	8.81	8.69	8.57	8.45	8.33	8.20	8.08	7.96
21°C	9.11	9.00	8.88	8.76	8.64	8.52	8.40	8.28	8.17	8.05	7.93	7.81
22°C	8.94	8.83	8.71	8.59	8.48	8.36	8.25	8.13	8.01	7.90	7.78	7.67
23°C	8.78	8.66	8.55	8.44	8.32	8.21	8.09	7.98	7.87	7.75	7.64	7.52
24°C	8.62	8.51	8.40	8.28	8.17	8.06	7.95	7.84	7.72	7.61	7.50	7.39
25°C	8.47	8.36	8.25	8.14	8.03	7.92	7.81	7.70	7.59	7.48	7.37	7.26
26°C	8.32	8.21	8.10	7.99	7.89	7.78	7.67	7.56	7.45	7.35	7.24	7.13
27°C	8.17	8.07	7.96	7.86	7.75	7.64	7.54	7.43	7.33	7.22	7.11	7.01
28°C	8.04	7.93	7.83	7.72	7.62	7.51	7.41	7.30	7.20	7.10	6.99	6.89
29°C	7.90	7.80	7.69	7.59	7.49	7.39	7.28	7.18	7.08	6.98	6.87	6.77
30°C	7.77	7.67	7.57	7.47	7.36	7.26	7.16	7.06	6.96	6.86	6.76	6.66
31°C	7.64	7.54	7.44	7.34	7.24	7.14	7.04	6.94	6.85	6.75	6.65	6.55

Table 4: Approximate Barometric Pressure at Different Elevations					
Elevation (feet)	**Pressure (mm Hg)**	**Elevation (feet)**	**Pressure (mm Hg)**	**Elevation (feet)**	**Pressure (mm Hg)**
0	760	2000	708	4000	659
250	753	2250	702	4250	653
500	746	2500	695	4500	647
750	739	2750	689	4750	641
1000	733	3000	683	5000	635
1250	727	3250	677	5250	629
1500	720	3500	671	5500	624
1750	714	3750	665	5750	618

DATA & CALCULATIONS

Dissolved Oxygen

Stream or lake: ______________________ Time of day: ______________________

Site name: ______________________ Student name: ______________________

Site number: ______________________ Student name: ______________________

Date: ______________________ Student name: ______________________

Column	A	B	C	D	E
Reading	Dissolved oxygen (mg/L)	Water temperature (°C)	Atmospheric pressure (mmHg)	100% dissolved oxygen (mg/L)	Percent saturation (%)
Example	8.2 mg/L	18.4°C	760 mmHg	9.5 mg/L	86 %
1					
2					

Average %	

Column Procedure:

A. Record the dissolved oxygen reading from computer.
B. Record the water temperature from a Temperature Probe or thermometer (Test 1).
C. Record the atmospheric pressure from a barometer or by using known altitude (see Table 4).
D. From Table 3, record the 100% dissolved oxygen value using measured temperature and atmospheric pressure.
E. Percent saturation = A / D X 100

Field Observations (e.g., weather, geography, vegetation along stream)______________________

____

____

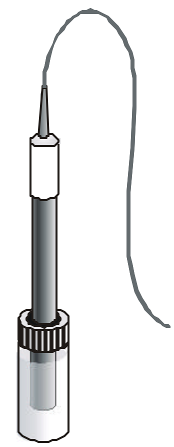
____

Test Completed: ______________ Date: ______

ADDITIONAL INFORMATION

Tips for Instructors

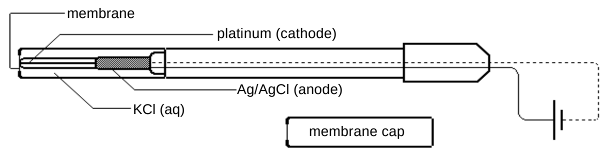
1. Before calibrating or taking measurements with the Dissolved Oxygen Probe, it is necessary to polarize the probe. Polarization is a 30-minute period required to warm up the probe. Think of it like a clothes iron that has to warm up before you can use it to iron your clothing. You must also keep the probe plugged in until you are finished taking all of your measurements. When the probe is unplugged from an active interface, it begins to cool down just like the clothes iron. It is a good idea to always start with fresh batteries or a full charge, if the unit is rechargeable, in the Vernier computer interface.



2. The best way to transport the Dissolved Oxygen Probe is to have it set up and polarizing while you are en route to the sampling site. Once the membrane cap of the probe is filled with electrode filling solution, you need to avoid letting the membrane dry out. To prevent drying, fill the DO calibration bottle with distilled water, fit the probe down into the lid, and tighten the lid onto the bottle. The probe tip should be submerged in the distilled water until you need to take samples.
3. To polarize the Dissolved Oxygen Probe, you need to have it plugged into an active interface. Make sure the interface box power is on and plug the DO probe into one of the interface ports. In 30 minutes, the DO probe will be polarized and ready to use.
4. When calibrating the Dissolved Oxygen Probe, it is important to be patient and permit the readings to stabilize.
5. As the Dissolved Oxygen Probe measures dissolved oxygen, it removes O_2 from the water sample at the junction of the probe membrane. If you leave the probe in one spot in the water sample, you will see your dissolved oxygen readings drop. To prevent this, it is important that students stir the probe gently and slowly through the sample as they take readings.
6. The gas-permeable plastic membrane on the Dissolved Oxygen Probe can become clogged by dirt and oil over time. Advise students to avoid touching the membrane at any time. If the water being sampled is murky or dirty, rinse the probe tip with distilled water after each use.
7. The electrode of the Dissolved Oxygen Probe is water tight and will not be damaged by water. The junction at the top of the electrode where the cable enters is not water tight and should not be submerged in water for any period of time. To take dissolved oxygen readings at various depths, use a Water Depth Sampler (order code WDS, $57). This device can be lowered to any desired depth and triggered to collect a representative water sample.
8. Step 6 of the student procedure provides several alternatives for loading or performing a Dissolved Oxygen Probe calibration:
   - The easiest option is to use the calibration that is stored with the Logger *Pro* experiment file, “Test 05 Dissolved Oxygen”.
   - Another option is to perform a two-point calibration in the lab, by choosing Calibrate from the Experiment menu (described in Steps 6a–i of the student procedure). After this calibration is completed, save the calibration values along with the experiment file. When students are ready to use the Dissolved Oxygen Probe, they can simply open the saved experiment file, and the new calibration will be loaded as well. Since there is very little change in Dissolved Oxygen Probe performance over short periods of time, we think this is a good way for students to handle dissolved-oxygen calibrations.
   - The third option is to have students perform the two-point calibration described in Steps 6a–i of the student procedure.

How the Dissolved Oxygen Probe Works

The Vernier Dissolved Oxygen Probe is a Clark-type polarographic electrode that senses the oxygen concentration in water and aqueous solutions. A platinum cathode and a silver/silver chloride reference anode in KCl electrolyte are separated from the sample by a gas-permeable plastic membrane.



A fixed voltage is applied to the platinum electrode. As oxygen diffuses through the membrane to the cathode, it is reduced:

$$\frac{1}{2} O_2 + H_2O + 2e^- \longrightarrow 2\ OH^-$$

The oxidation taking place at the reference electrode (anode) is:

$$Ag + Cl^- \longrightarrow AgCl + e^-$$

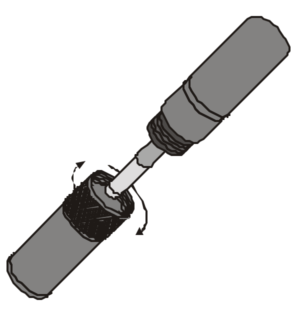
Accordingly, a current will flow that is proportional to the rate of diffusion of oxygen, and in turn to the concentration of dissolved oxygen in the sample. This current is converted to a proportional voltage, which is amplified and read by any of the Vernier lab interfaces.

Preparing the Dissolved Oxygen Probe for Use

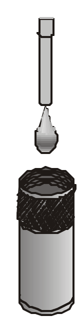
Read this procedure carefully before using your Dissolved Oxygen Probe. Following the directions outlined in this section will help ensure that you will obtain accurate dissolved oxygen readings with the probe.

1. Remove the blue protective cap from the tip of the probe. This protective cap can be discarded once the probe is unpacked.

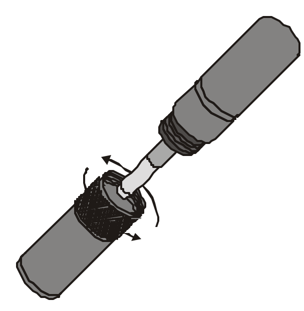
2. Unscrew the membrane cap from the tip of the electrode body with a counterclockwise turn. **Important**: Do not touch the membrane at the very tip of the probe. Contact with the membrane could result in damage to the membrane surface and improper operation of the probe. Using the pipet supplied, fill the membrane cap with approximately 1 mL of DO Electrode Filling Solution (included). Then carefully thread the membrane cap (clockwise) onto the electrode body, being careful not to over tighten the cap. Rinse the electrode with distilled water and carefully wipe it dry with a lab wipe. Again, take care not to touch the membrane itself.



Remove membrane cap



Add electrode filling solution



Replace membrane cap

3. Place the probe in a 250-mL beaker containing about 75 mL of distilled water. Connect the probe to your interface.

4. With the probe connected to an *active* interface, allow it to polarize for about 30 minutes.

5. After polarization has taken place for 30 minutes, the Dissolved Oxygen Probe is ready to take readings or to be calibrated.

Storage of the Dissolved Oxygen Probe

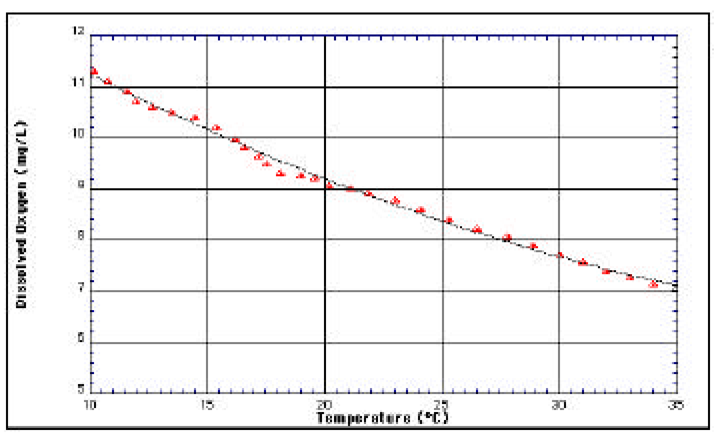
Follow these steps when storing the electrode:

- **Long-term storage (more than 24 hours):** Remove the membrane cap and rinse the inside and outside of the cap with distilled water. Shake the membrane cap dry. Also rinse and dry the exposed anode and cathode inner elements (blot dry with a lab wipe). Reinstall the membrane cap loosely onto the electrode body for storage. Do not screw it on tightly.
- **Short-term storage (less than 24 hours):** Store the Dissolved Oxygen Probe with the membrane end submerged in about 1 inch of distilled water.

Automatic Temperature Compensation

Your Vernier Dissolved Oxygen Probe is automatically temperature compensated, using a thermistor built into the probe. The temperature output of this probe is used to automatically compensate for changes in permeability of the membrane with changing temperature. If the probe was not temperature compensated, you would notice a change in the dissolved oxygen reading as temperature changed, even if the actual concentration of dissolved oxygen in the solution did not change. Here are two examples of how automatic temperature compensation works:

- If you calibrate the Dissolved Oxygen Probe in the lab at 25°C and 760 mm Hg barometric pressure (assume salinity is negligible), the value you entered for the saturated oxygen calibration point would be 8.36 mg/L (see Table 3). If you were to take a reading in distilled water that is saturated with oxygen by rapid stirring, you would get a reading of 8.36 mg/L.
  If the water sample is then cooled to 10°C with no additional stirring, the water would no longer be saturated (cold water can hold more dissolved oxygen than warm water). Therefore, the reading of the temperature-compensated Dissolved Oxygen Probe would still be 8.36 mg/L.
- If, however, the solution was cooled to 10°C *and* continually stirred so it remained saturated by dissolving additional oxygen, the temperature-compensated probe would give a reading of 11.35 mg/L—the value shown in Table 3. **Note:** Temperature compensation *does not mean* that the reading for a saturated solution will be the same at two different temperatures—the two solutions have different concentrations of dissolved oxygen, and the probe reading should reflect this difference.



Saturated Dissolved Oxygen vs. Temperature Data

Sampling in Ocean Salt

Water or Tidal Estuaries

(at salinity levels greater than 1000 mg/L)

Dissolved Oxygen concentration for air saturated water at various salinity values, $DO_{(salt)}$, can be calculated using the formula:

$$DO_{(salt)} = DO - (k \bullet S)$$

– $DO_{(salt)}$ is the concentration of dissolved oxygen (in mg/L) in salt-water solutions.

– *DO* is the dissolved oxygen concentration for air-saturated distilled water as determined from Table 3.

– *S* is the salinity value (in ppt). Salinity values can be determined using the Vernier Chloride Ion-Selective Electrode or Conductivity Probe as described in Test 15.

– *k* is a constant. The value of *k* varies according to the sample temperature, and can be determined from Table 5.

Table 5: Salinity Correction Constant Values

Temp. (°C)	Constant, k	Temp. (°C)	Constant, k	Temp. (°C)	Constant, k	Temp. (°C)	Constant, k
1	0.08796	8	0.06916	15	0.05602	22	0.04754
2	0.08485	9	0.06697	16	0.05456	23	0.04662
3	0.08184	10	0.06478	17	0.05328	24	0.04580
4	0.07911	11	0.06286	18	0.05201	25	0.04498
5	0.07646	12	0.06104	19	0.05073	26	0.04425
6	0.07391	13	0.05931	20	0.04964	27	0.04361
7	0.07135	14	0.05757	21	0.04854	28	0.04296

Example: Determine the saturated DO calibration value at a temperature of 23°C and a pressure of 750 mm Hg, when the Dissolved Oxygen Probe is used in seawater with a salinity value of 35.0 ppt.

First, find the dissolved oxygen value in Table 3 (DO = 8.55 mg/L). Then find *k* in Table 5 at 23°C (k = 0.04662). Substitute these values, as well as the salinity value, into the previous equation:

$$DO_{(salt)} = DO - (k \bullet S) = 8.55 - (0.04662 \times 35.0) = 8.55 - 1.63 = 6.92 \text{ mg/L}$$

Use the value 8.46 mg/L when performing the saturated DO calibration point (water-saturated air), as described in Step 6. The Dissolved Oxygen Probe will now be calibrated to give correct DO readings in salt-water samples with a salinity of 35.0 ppt.

Important: For most dissolved oxygen testing, it is *not* necessary to compensate for salinity; for example, if the salinity value is 0.5 ppt, using 25°C and 760 mm Hg, the calculation for DO(s) would be:

$$DO_{(salt)} = DO - (k \bullet S) = 8.36 - (0.04498 \times 0.5) = 8.36 - 0.023 = 8.34 \text{ mg/L}$$

At salinity levels less than 1.0 ppt, neglecting this correction results in an error of less than 0.2%.